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## Spectroscopic Studies of Mixed Phosphoric-Carboxylic Imides

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# SPECTROSCOPIC STUDIES OF MIXED PHOSPHORIC-CARBOXYLIC IMIDES

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Abstract IR spectroscopic studies of hydrogen bonding (intra-vs intermolecular) in selected mixed phosphoric-carboxylic imides (RO)  $_2P(O)NR'C(O)R"$  1, R = Et, R' = H, alkyl, R" = OEt, Ph, are presented.

#### INTRODUCTION

IR spectroscopy proved ideal to determine the position of the double bonds and possible intramolecular hydrogen bond formation<sup>1</sup> in the title compounds. Compounds 1 and their salts show antiviral activity<sup>2</sup> and also some interesting complexing properties.<sup>3</sup> These systems exhibit ambident properties both in electrophilic and nucleophilic reactions.<sup>4</sup>

### RESULTS AND DISCUSSION

While  $\beta$ -dicarbonyl compounds exist as an intramolecularly hydrogen bonded keto-enol system, no intramolecular hydrogen bond formation was observed for  $\beta$ -diphosphonates. For mixed methylene derivatives, the C=O group exists in its enolic form, while the P=O group forms intermolecular hydrogen bonds with external donors. Methanol proved to be a too weak hydrogen donor towards substrates 1, while phenol caused a strong to modest  $v_{C=O}$  shift indicating that the C=O group is a better hydrogen bonding acceptor than the P=O group. The large P=O frequency shift difference for the *N*-unsubstituted and *N*-substituted ethoxycarbonyl phosphoramidate substrates in the absence of an external donor, indicates the existence of tautomerism towards the P=O group and not the C=O group.

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